



## **EFFECTIVE: SEPTEMBER 2004**

### **CURRICULUM GUIDELINES**

**A.** Division: Science and Technology

Effective Date:

September 2004

**B.** Department / Program  
Area: Chemistry

Revision

X

h descriptor)

6

Number of Weeks per Semester: 15

**K:** Maximum Class Size:

36

**L:** PLEASE INDICATE:

Non-Credit

**M:** Course Objectives / Learning Outcomes

Upon completion of this course, the students will:

1. Carry out measurements using the correct number of significant figures, and express the precision using absolute or relative uncertainties.
2. Given a set of experimental data, calculate the average value, the average deviation, and the standard deviation.
3. Solve stoichiometry problems of the following types: percentage composition/empirical formula, gram-gram or gram-volume (of a gas), solution stoichiometry.

molecules; Valence Bond Theory: hybridization, orbital diagrams; Molecular Orbital Theory: shapes and energies of molecular orbitals, bond order, intermolecular forces, and hydrogen bonding.

<b>R:</b> Prior Learning Assessment and Recognition: specify whether course is open for PLAR
No

\_\_\_\_\_  
Course Designer(s)

\_\_\_\_\_  
Education Council / Curriculum Committee Representative

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Dean / Director

\_\_\_\_\_  
Registrar